**Reflection**

03/05/2019

1. **What have I learned in this unit?**

In this unit of Chemical bonding and structure, we basically learned intramolecular forces and intermolecular forces. We first learned the difference between ionic and covalent bonding, along with the classification of metals and anion, and the naming rules for each of the bonding. Then we proceed to learn to draw and interpret Lewis diagrams of atoms, including the concept of lone pairs and bonding electrons, and skeleton structure for molecules. Afterwards, we learned parity, and VSEPR to gain a more insightful understanding for intramolecular forces. Finally, we explored different types of intermolecular forces and the reason behind their difference.

1. **How will each of the topics learned help you with the project? How do you apply this to MSA?**

Both intramolecular and intermolecular forces are important to my project in order to make molecules or atoms to self-assemble, because molecule its self can also be seen as a process of self-assembly. By studying the formation of molecules and substances, I now know what can cause the attraction between molecules and atoms, and which one is the strongest and which ones are weaker. I can apply that knowledge of different attractions to molecular self-assembly, as thus better control the self-assembly process.

1. **What do you need clarified in order to complete the project?**

How many different ways do scientists have nowadays in order to control molecular self-assembly?

Is there really a way for us to control, i.e. start and pause immediately, the process of molecular self-assembly? What can we do?

What other conditions molecules need to meet besides the attraction forces and appropriate temperature?